

## *Chapter 10 Chemical Quantities Guided Practice Answers*







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Guided Notes Chapter 10: Chemical Quantities Name\_\_\_\_\_ 10.1 The Mole: A Measurement of Matter  
1. What is the mole? 2. What is Avogadro's number? 3. What kinds of things can you count using  
the mole? 4. What are three ways of measuring the amount of a substance? 5.

### Guided Notes Chapter 10: Chemical Quantities Name 10.1 The ...

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of  
a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$   
particles of that substance • The mole was founded by a scientist named Avogadro, and he decided  
to use the

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municipalities this chapter is intended to shed light on some key strategic concepts developed later  
in this book including the fluidity of firm boundaries the importance of scale and 5 / 6.

### Chapter 10 Chemical Quantities Guided Practice Answers

Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT ... 92 Guided  
Reading and Study Workbook CHAPTER 10,Chemical Quantities(continued) 9. Complete the table  
about representative particles and moles. The Mass of a Mole of an Element (pages 293-294) 10.

### SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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Hs. Chapter 10 CHEMICAL QUANTITIES The MOLE Avogadros Hypothesis Equal volumes of gases (@  
same T and p)

**Chapter 10 Packet Answer Key - Chapter 10 CHEMICAL ...**

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

**Chemical Quantities - Weebly**

10) What is the molarity of solution made by dissolving 10 g of NaCl into enough water to make 250 mL of saltwater solution? or M L mol L mL g NaCl mol NaCl mL solution g NaCl 0.68 0.68 1 1000 58.44 1 250 10 u u III. Percent Composition and Chemical Formulas 11) A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g

**Chemistry--Chapter 7: Chemical Quantities**

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